

Chemical Equations And Reactions Test Review Answers

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Chemical Equations And Reactions Test

Todd Helmenstine. Updated June 27, 2019. Chemical reactions have the same number of atoms before the reaction as after the reaction. Balancing chemical equations is a basic skill in chemistry and testing yourself helps retain important information. This collection of ten chemistry test questions will give you practice in how to balance chemical reactions .

Balancing Equations Chemistry Test Questions

In the given reaction, Zn is oxidized; H + is reduced . Question 7: Reaction which has equal number of atoms of all the elements on both sides of a chemical equation is called. equilibrium reaction. spontaneous reaction. balanced reaction. none of these. Correct Option is : 3. Solution :

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Reaction information is shown using word and symbol equations. The simplest ratio of atoms of each element in a compound is called the empirical formula. Mass is conserved in chemical reactions.

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Test , Class 10, Chemistry, CBSE- Chemical Reactions and Equations.

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As a result of the experiments of Lavoisier, you know that in a chemical reaction, the mass of the products _____ . always equals the mass of the reactants. In chemical equations, _____ represent the number of units of each substance. coefficients.

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Write values of a,b and c if following chemical reaction is balanced . $a\text{Mg} + b\text{O}_2 \rightarrow c\text{MgO}$. (i) $a=1, b=2, c=2$. (ii) $a=2, b=1, c=2$. (iii) $a=2, b=2, c=2$. (iv) $a=1, b=2, c=1$. Q10. Write values of a, b and c so that following chemical equation is balanced. $a\text{H}_2 + b\text{O}_2 \rightarrow c\text{H}_2\text{O}$.

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Question 1. Which of the following reactions involves the combination of two elements :-
(a) $\text{CaO} + \text{CO}_2 \rightarrow \text{CaCO}_3$ (b) $4\text{Na} + \text{O}_2 \rightarrow 2\text{Na}_2\text{O}$ (c) $\text{SO}_2 + (1/2)\text{O}_2 \rightarrow \text{SO}_3$ (d) $\text{NH}_3 + \text{HCl} \rightarrow \text{NH}_4\text{Cl}$

Question 2. When hydrogen sulphide gas is passed through a blue solution of copper sulphate, a black precipitate of copper sulphide is obtained and the sulphuric acid so formed remains in the solution.

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Chemical Reactions Test. acids, bases, indicators, pH, etc.... STUDY. PLAY. Matter. anything that has mass and takes up space. Chemistry. ... a number placed in front of a chemical formula in an equation. 4 Types of Chemical Reactions. synthesis, decomposition, single replacement and double replacement.

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$2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \longrightarrow 2\text{H}_2\text{O}(\text{l})$ (ii) Condition in which reaction takes place are written on the arrow head, e.g. Consider the following chemical reaction. $\text{X} + \text{Barium chloride} \longrightarrow \text{Y} + \text{Sodium chloride}$. (White ppt) (a) Identify 'X' and 'Y'. (b) The type of reaction. (a) 'X' is Na_2SO_4 and Y is BaSO_4 .

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Test # Observations Balanced Equation Physical or Chemical Reaction Type Splint Test For Double

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An aqueous solution is a solution in which the solvent is water. It is mostly shown in chemical equations by appending (aq) to the relevant chemical formula. For example, a solution of table salt, or sodium chloride (NaCl), in water would be represented as $\text{Na}^+ + \text{Cl}^-$ (aq). The word aqueous (which comes from aqua) means pertaining to, related to, similar to, or dissolved in, water.

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