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is produced by the following reaction.

$$3\text{Si(s)} + 2\text{N}_2\text{(g)} \rightarrow \text{Si}_3\text{N}_4\text{(s)}$$

a. Chapter 12

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In Example 12.2.1 and Example 12.2.2, the identity of the limiting reactant has been

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apparent: $[\text{Au}(\text{CN})_2]^-$, LaCl_3 , ethanol, and para-nitrophenol. When the limiting reactant is not apparent, we can determine which reactant is limiting by comparing the molar amounts of the reactants with their coefficients in the balanced chemical equation, just as we did in Section 11.4 .

Chapter 12.2: **Stoichiometry of**

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**Reactions in
Solution ...**

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In any stoichiometry
problem, the first step
is always to calculate
the number of moles of
each reactant present.
In this case, we are
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given the mass of K₂

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engines and turbines. It

is produced by the fol-

lowing reaction. $3\text{Si}(s)$

$2\text{N}_2(g) \rightarrow \text{Si}_3\text{N}_4(s)$

a. Interpret the

equation in terms of

particles, moles, and

masses. b.

Stoichiometry - MARRIC

A In any stoichiometry

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problem, the first step
is always to calculate
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Data AnalysisData

Analysis 1. A sample of

aluminum is placed in

a 25-mL graduated

cylinder containing

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10.0 mL of water. The level of water rises to 18.0 mL. Aluminum has a density of 2.7 g/mL. Calculate the mass of the sample.

2. Saturn is about 1 429 000 km from the Sun.

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Stoichiometry
Problems. Directions:

Solve each of the following problems.

Assume excess of any reactant that isn't mentioned, unless otherwise . specified.

Show your work, including proper units, to earn full credit. 1.

Given the balanced equation, show what the following molar ratios (i.e., the conversion factors ...

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Stoichiometry Practice Problems Answer Key A

In any stoichiometry problem, the first step is always to calculate the number of moles

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Stoichiometry study guide by Leahrosner includes 30 questions covering vocabulary, terms and more.

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